

SrNi₃(P₂O₇)₂**Michael Bolte**

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Key indicators

Single-crystal X-ray study

$T = 173$ K

Mean $\sigma(\text{P}-\text{O}) = 0.001$ Å

R factor = 0.021

wR factor = 0.054

Data-to-parameter ratio = 20.0

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

The title compound, strontium trinickel bis(diphosphate), SrNi₃(P₂O₇)₂, is isostructural with all other mixed metal pyrophosphates having the same stoichiometry. The Sr and one of the Ni ions are located on a crystallographic centre of inversion.

Comment

Pyrophosphates are of interest because of their complex network architecture and several structures have been previously determined. The title compound, SrNi₃(P₂O₇)₂ (Fig. 1), has already been investigated by powder diffraction (El-Bali, 1993), but since its single-crystal structure has not been determined yet, it is presented here.

The Sr and one of the Ni ions are located on a crystallographic centre of inversion. All other atoms occupy general positions. The structure consists of infinite zigzag chains of NiO₆ octahedra sharing either *trans* or skew edges. These chains are connected by P₂O₇ moieties to form a three-dimensional network into which the Sr ions are incorporated *via* connections to eight O atoms (Table 1). The structure of SrNi₃(P₂O₇)₂ is isostructural with all other metal pyrophosphates having the same stoichiometry: Ni₃Pb(P₂O₇)₂ (Krasnikov *et al.*, 1985), CaNi₃(P₂O₇)₂, CaCo₃(P₂O₇)₂ and SrFe₃(P₂O₇)₂ (Lii *et al.*, 1993), and PbCo₃(P₂O₇)₂ and PbFe₃(P₂O₇)₂ (Elmarzouki *et al.*, 1995).

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Experimental

A mixture of (NH₄)₂HPO₄, Sr(OH)₂ and NiCl₂ was ground together and heated to approximately 800 K. The molten mass was maintained at this temperature for two days and then cooled down to room temperature (El-Bali, 1993).

Crystal data

SrNi₃(P₂O₇)₂

$M_r = 611.63$

Monoclinic, $P2_1/c$

$a = 7.4116$ (4) Å

$b = 7.6542$ (3) Å

$c = 9.4486$ (6) Å

$\beta = 112.194$ (5)°

$V = 496.30$ (5) Å³

$Z = 2$

$D_x = 4.093$ Mg m⁻³

Mo $K\alpha$ radiation

Cell parameters from 39700 reflections

$\theta = 3.8$ – 37.3 °

$\mu = 11.69$ mm⁻¹

$T = 173$ (2) K

Block, yellow

$0.24 \times 0.18 \times 0.16$ mm

Data collection

Stoe IPDS II two-circle diffractometer

ω scans

Absorption correction: empirical (*MULABS*; Spek, 1990; Blessing, 1995)

$T_{\min} = 0.101$, $T_{\max} = 0.156$

26 742 measured reflections

2076 independent reflections

2027 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.052$

$\theta_{\max} = 34.3$ °

$h = -11 \rightarrow 11$

$k = -12 \rightarrow 10$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.021$
 $wR(F^2) = 0.054$
 $S = 1.13$
 2076 reflections
 104 parameters

$w = 1/[\sigma^2(F_o^2) + (0.0304P)^2 + 0.4865P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.78 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.78 \text{ e } \text{\AA}^{-3}$
 Extinction correction: *SHELXL97*
 Extinction coefficient: 0.096 (2)

Table 1

Selected bond lengths (Å).

Sr1—O7	2.5453 (12)	Ni1—O2 ^{vi}	2.0841 (12)
Sr1—O2 ⁱ	2.6234 (13)	Ni1—O1	2.0901 (13)
Sr1—O3 ⁱⁱ	2.6541 (13)	Ni1—O3 ⁱⁱ	2.1961 (12)
Sr1—O1 ⁱⁱⁱ	2.7169 (13)	Ni2—O6 ⁱⁱ	2.0458 (12)
Ni1—O6 ^{iv}	2.0415 (12)	Ni2—O7 ^{vii}	2.0552 (12)
Ni1—O5 ^v	2.0606 (12)	Ni2—O3 ⁱⁱ	2.1163 (12)

Symmetry codes: (i) $x, 1+y, z$; (ii) $1-x, \frac{1}{2}+y, \frac{1}{2}-z$; (iii) $x, \frac{3}{2}-y, z-\frac{1}{2}$; (iv) $x-1, \frac{3}{2}-y, z-\frac{1}{2}$; (v) $x-1, y, z$; (vi) $1-x, 1-y, -z$; (vii) $1-x, 2-y, -z$.

Data collection: *X-AREA* (Stoe & Cie, 2001); cell refinement: *X-AREA*; data reduction: *X-AREA*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1990); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *XP* in *SHELXTL-Plus* (Sheldrick, 1991).

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References

- Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
 El-Bali, B. (1993). PhD thesis, Université Mohammed V, Rabat, Morocco.

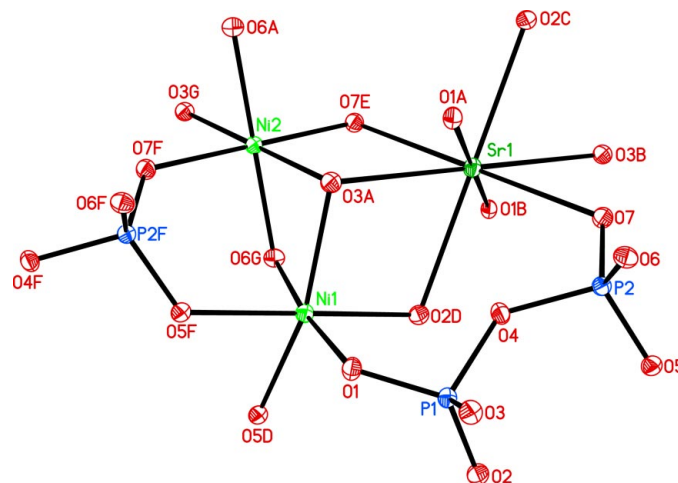


Figure 1

A perspective view of the title compound with the atom-numbering scheme. Displacement ellipsoids are at the 50% probability level. Symmetry operators: (A) $1-x, y+\frac{1}{2}, \frac{1}{2}-z$; (B) $x, \frac{3}{2}-y, z-\frac{1}{2}$; (C) $x, y+1, z$; (D) $1-x, 1-y, z$; (E) $1-x, 2-y, -z$; (F) $x-1, y, z$; (G) $x-1, \frac{3}{2}-y, \frac{1}{2}-z$.

- Elmarzouki, A., Boukhari, A., Berrada, A. & Holt, E. M. (1995). *J. Solid State Chem.* **118**, 202–205.
 Krasnikov, V. V., Konstant, Z. A. & Bel'skij, V. K. (1985). *Izv. Akad. Nauk SSSR Neorg. Mater.* **9**, 1560–1563.
 Lii, K.-H., Shih, P.-F. & Chen, T.-M. (1993). *Inorg. Chem.* **32**, 4373–4377.
 Sheldrick, G. M. (1990). *Acta Cryst.* **A46**, 467–473.
 Sheldrick, G. M. (1991). *SHELXTL-Plus*. Release 4.1. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
 Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
 Spek, A. L. (1990). *Acta Cryst.* **A46**, C-34.
 Stoe & Cie (2001). *X-AREA*. Stoe & Cie, Darmstadt, Germany.